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Example problems: Rhoda Tiller looks up to see a flowerpot fall off a ledge 9.2 m above her head.

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Aditi Acharya Bohr's Energy Levels •

Electrons in certain energy levels •

Lower energy levels: closer to nucleus •

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Higher energy levels: further from nucleus • Ground State: electron in lowest energy level possible Excited Atom • Atom has absorbed energy • Excited state is unstable • Atom soon emits same amount of energy absorbed • energy seen as visible light Wave Description of Light Wavelength (λ): distance between corresponding points

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levels- farther from nucleus
4. ground state- electron in lowest energy level possible.

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Season 1 Episode 303 | 22m 11s Atomic

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Theory: Electrons on the Atom Bohr's model of the atom are described. Students calculate wavelength, frequency, or the velocity of a wave using the wave ...

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